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National MDCAT 2020 Results

190+

35 Students

185+

218 Students

180+

677 Students



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National MDCAT 2020 Results

190+

35 Students

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Education



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Regards.Huzaiifa Saeed,Usama Sohail

Fizza Maryum,Javeria Shukor

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QUIZZES

Practice Test-1(Introduction
to Fundamental Concepts of Chem...



10 Questions



7 min

Topics

Atomic mass

Start Quiz

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06 : 36



2/10



7 min



Hint

Q : Which of the following term is used for the mass of chlorine 35.5 amu

A

Relative atomic mass

B

Relative Molecular mass

C

Mass number

D

Relative isotopic mass

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06 : 34



3/10



7 min



Hint

Q : Which of the following statement is wrong about isotopes

A

They possess different mass number

B

They possess same chemical properties

C

They possess different physical properties

D

They possess different position in the periodic table

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06 : 31



4/10



7 min



Hint

Q : Boron has two stable isotopes, ^{10}B (19%) and ^{11}B (81%) Find the average atomic mass of Boron

A

$$\frac{(19 \times 10) + (81 \times 10)}{100}$$

B

$$\frac{(19 \times 10) + (81 \times 11)}{100}$$

C

$$\frac{(81 \times 10) + (19 \times 11)}{100}$$

D

$$\frac{(19 \times 10) \times (81 \times 11)}{100}$$

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06 : 29



5/10



7 min



Hint

Q : In nature the ratio of relative percentage abundance of the isotopes ^{35}Cl and ^{37}Cl is _____ respectively



1 : 3



3 : 1



1 : 1



1 : 4

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06 : 27



6/10



7 min



Hint

Q : 1 amu is equal to

A

$1.661 \times 10^{-27} \text{ kg}$

B

$1.661 \times 10^{-24} \text{ g}$

C

$1.661 \times 10^{-21} \text{ mg}$

D

All of these

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06 : 24



7/10



7 min



Hint

Q : The atomic weight of B is 10.8. There are only two naturally occurring isotopes of boron and . The natural abundance of the isotope must be



10%



20%



50%



80%

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06 : 21



8/10



7 min



Hint

Q : Total number of radioactive isotopes produced through artificial disintegration



240



40



300



340

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06 : 18



9/10



7 min



Hint

Q : Which pair of elements have same number of isotopes



B, Cl



C, I



Na, Ca



Ni, Cd

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06 : 16



10/10



7 min



Hint

Q :

The atomic weight of Cu is 63.546. There are only two naturally occurring isotopes of copper, ^{63}Cu and ^{65}Cu . The natural abundance of the ^{63}Cu isotope must be

A

10%

B

30%

C

50%

D

70%

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QUIZZES

Practice Test-2 (Introduction
to Fundamental Concepts of Chem...



10 Questions



7 min

Topics

Concept of mole

Start Quiz

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1/10



7 min



Hint

Q : One mole of Hydrogen and oxygen have same at STP

A

Gram molecular weight

B

Protons in the molecules

C

molar volume

D

Electrons in the valence shell

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1

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2/10



7 min



Hint

Q : 18.0g of glucose contains number of hydrogen atoms

A

7.2×10^{23}

B

6.3×10^{23}

C

2.7×10^{23}

D

3.6×10^{23}

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3/10



7 min



Hint

Q : One mole of H_2SO_4 contain

A

6.022×10^{23} number of hydrogen atoms

B

3.011×10^{23} atoms of Sulphur

C

2.4088×10^{24} atoms of oxygen

D

2.4088×10^{24} molecules of oxygen

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1

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7



4/10



7 min



Hint

Q : 4g H_2 reacts with 32.0g O_2 to produce water. Which of the following statements is correct

A

H_2 -limiting reactant

B

O_2 -non-limiting reactant

C

2.0 mole water is produced

D

1 mole water is produced

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5/10



7 min



Hint

Q : 2.8g of N_2 molecule contains number of chemical bonds

A

6.02×10^{22}

B

1.204×10^{23}

C

1.8×10^{23}

D

1.8×10^{22}

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6/10



7 min



Hint

Q : The number of moles of KMnO_4 that contain 1 mole of oxygen

A

2 moles

B

0.5 moles

C

0.25 moles

D

1.5 moles

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7/10



7 min



Hint

Q : The number of atoms in one gram atom of an element is

A

N_A of atoms

B

N_A of ions

C

N_A of molecules

D

N_A of formula unit

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8/10



7 min



Hint

Q : One gram molecule of different gases have all the following properties same at STP except

A

Molecules

B

Moles

C

Volume

D

Masses

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9/10



7 min



Hint

Q : The total number of O-atoms in 18g of glucose are

A

6.02×10^{22}

B

6.02×10^{23}

C

3.6×10^{23}

D

3.6×10^{22}

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10/10



7 min



Hint

Q : How many moles of neutron are present in one mole of heavy water



10



18



8



20

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QUIZZES

Practice Test-3(Introduction
to Fundamental Concepts of Chem...



10 Questions



7 min

Topics

Avogadro's number and Vm

Start Quiz

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06 : 58



1/10



7 min



Hint

Q : The volume occupied by 1.6g of O_2 at STP is

A

2.24dm³

B

22.4dm³

C

1.12dm³

D

112cm³

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06 : 56



2/10



7 min



Hint

Q : Eight grams of methane occupies volume at STP

A

22.4 dm³

B

2.24 dm³

C

1.12 dm³

D

11.2 dm³

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06 : 49



3/10



7 min



Hint

Q : Avogadro's number is the number of molecules present in

A

1 dm³ of molecule

B

1 g of formula mass

C

Gram molecular mass

D

1 g of atom

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06 : 47



4/10



7 min



Hint

Q :

15 gram of a gas occupies 11.2 dm^3 at S.T.P,
the gas is



CO



NO



CO₂



N₂O

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06 : 45



5/10



7 min



Hint

Q : 4 g of CH_4 gas has molar volume at S.T.P

A

22.414 cm^3

B

5.60 dm^3

C

11.2 cm^3

D

22414 cm^3

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06 : 43



6/10



7 min



Hint

Q : Hydrogen and oxygen have same at STP

A

Gram molecular weight

B

Protons in the molecules

C

Gram molecular volume

D

Electrons in the valence shell

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06 : 41



7/10



7 min



Hint

Q : Equal volumes of N_2O and CO_2 are taken in identical conditions, the correct relation between the masses of two gases is



$\text{N}_2\text{O} > \text{CO}_2$



$\text{N}_2\text{O} < \text{CO}_2$



$\text{N}_2\text{O} = \text{CO}_2$



$\text{N}_2\text{O}^3 \text{CO}_2$

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06 : 38



8/10



7 min



Hint

Q : Which one is incorrect relation at STP

A

6g of carbon = 3.01×10^{23} atoms

B

11.2 dm³ of CO₂ = 3.01×10^{23} molecules

C

49 g of H₂SO₄ = 4 moles of atoms

D

1 mole of sucrose = 45 moles of atoms

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06 : 34



9/10



7 min



Hint

Q : Avogadro's number is the number of molecules present in

A

1 dm³ of molecule

B

1 g of hydrogen gas

C

1 g atom

D

Gram molecular mass

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4

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10

06 : 32



10/10



7 min



Hint

Q : The volume of oxygen gas is 1.12dm^3 at STP , the mass of oxygen approximately is



3.2g



1.6g



2.8g



1.4g

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QUIZZES

Practice Test-4 (Introduction
to Fundamental Concepts of Chem...



10 Questions



7 min

Topics

Empirical and Molecular formulae

Start Quiz

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Q : Which of the following term is not used for ionic compounds?

A

Formula unit

B

Empirical formula

C

Molecular formula

D

Formula Mass

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2/10



7 min



Hint

Q : Which one of the following statements is not involved in the determination of empirical formula

A

%age of each element

B

gram atom of each element

C

isotopes of each element

D

atomic ratio of element

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Q : Which is correct statement

A

Rectified spirit is 100% ethanol

B

Molar volume of an ideal gas is 24 dm³ at STP

C

R_f value for red ink has units of cm⁻¹

D

Quantitative analysis involves four steps

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4/10



7 min



Hint

Q : Elemental analysis is performed to determine

A

Molar mass of the compound

B

Structural formula of a compound

C

Empirical formula of a compound

D

Mass of halogen present in a compound

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1

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5/10



7 min



Hint

Q :

There are different steps in determining the empirical formula

Step I. Calculating the number of gram atom

Step II. Determining the atomic ratio

Step III. Determining the percentage composition

What is the correct sequence of the above steps?

A

I, II, III

B

III, II, I

C

II, I, III

D

III, I, II

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6/10



7 min



Hint

Q : The simplest formula of a compound containing 50% of element X (At.wt = 10) and 50% of element Y (At. wt = 20) is

A

XY

B

XY₂

C

X₂Y

D

X₂Y₃

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7/10



7 min



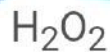
Hint

Q : While determine molecular formula, the simple multiple 'n' is not unity for

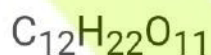
A



B



C



D



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1

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8/10



7 min



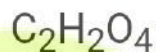
Hint

Q : An acid with molecular mass 104 contains 34.6% C, 3.85% H and rest is O. The molecular formula of acid is

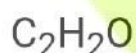
A



B



C



D



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Q : A compound contains 50% S and 50% O by mass. The empirical formula of compound is

A



B



C



D



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10/10



7 min



Hint

Q : A pair of compounds that has same empirical formula

A

Acetic acid and glucose

B

Acetic acid and formic acid

C

Formic acid and sucrose

D

Ethane and Ethyne

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QUIZZES

Practice Test-5 (Introduction
to Fundamental Concepts of Chem...



10 Questions



7 min

Topics

Stoichiometry

Start Quiz

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06 : 57



1/10



7 min



Hint

Q : For a 10% solution of acetic acid in water contain 1.850 mole is the calculate value of

A

Normality

B

Molarity

C

Molality

D

Formality

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06 : 54



2/10



7 min



Hint

Q : A solution of glucose is 10% W/V. The volume in which 1g mole of it is dissolved will be

A

1dm^3

B

1.8dm^3

C

200cm^3

D

900cm^3

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06 : 52



3/10



7 min



Hint

Q : A solution of 92g of ethanol, 96g methanol and 90g water has mole fraction of ethanol equal to

A

0.1

B

0.4

C

0.2

D

0.5

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06 : 50



4/10



7 min



Hint

Q : Which of the following solution pairs can be separated into its pure components by fractional distillation?

A

Benzene — toluene

B

Water -HNO₃

C

Water-HCl

D

Water-C₂H₅OH

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06 : 48



5/10



7 min



Hint

Q : If 5.85 of NaCl are dissolved in 90g of water, the mole fraction of solute is

A

0.01

B

0.025

C

0.0196

D

0.0382

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06 : 46



6/10



7 min



Hint

Q : If 5.85 g of NaCl are dissolved in 90g of water, the mole fraction of NaCl is:

A

0.1

B

0.01

C

0.2

D

0.0196

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06 : 44



7/10



7 min



Hint

Q : The amount of NaOH required to make 1dm^3 of 0.5M aqueous solution



2g



10g



4g



20g

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06 : 41



8/10



7 min



Hint

Q : Number of moles of solute present in 50cm^3 of 0.1M aqueous solution is

A

0.05

B

0.005

C

0.001

D

0.1

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06 : 39



9/10



7 min



Hint

Q : An aqueous solution of 0.05M NaOH is available. What is %age W/V of this solution.

A

2%

B

0.2%

C

4%

D

10%

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06 : 36



10/10



7 min



Hint

Q : A solution of glucose is 10% W/V. The volume in which 1g mole of it is dissolved will be

A

1dm^3

B

1.8dm^3

C

200cm^3

D

900cm^3

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Date: -/ /20

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Practice Test 1	Practice Test 2	Practice Test 3
1-B	1-C	1-C
2-A	2-A	2-D
3-D	3-C	3-C
4-B	4-C	4-B
5-B	5-C	5-B
6-D	6-C	6-C
7-D	7-A	7-C
8-C	8-D	8-C
9-A	9-C	9-D
10-D	10-A	10-B

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Practice Test 4 Practice Test 5 Practice Test 6

1-C	1-A	1-C
2-C	2-A	2-B
3-D	3-C	3-C
4-C	4-B	4-A
5-D	5-B	5-C
6-C	6-A	6-D
7-B	7-D	7-D
8-A	8-A	8-B
9-A	9-B	9-B
10-A	10-B	10-B

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QUIZZES

Test Unit-1(Introduction to
Fundamental Concepts of Chemist...



30 Questions



25 min

Topics

Atomic mass, Concept of mole, Avogadro's
number and V_m , Empirical and Molecular
formulae, Stoichiometry, Units of
Concentration

Start Quiz

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24 : 58



1/30



25 min



Hint

Q : Haemoglobin molecule is _____ times then heavier H_2

A

10,000

B

68,000

C

1000

D

34,000

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6

7

24 : 56



2/30



25 min



Hint

Q :

The isotopes $^{16}_8\text{O}$, $^{24}_{12}\text{Mg}$, $^{28}_{14}\text{Si}$, $^{40}_{20}\text{Ca}$, $^{56}_{26}\text{Fe}$ form nearly _____ of earth crust

A

20%

B

30%

C

40%

D

50%

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24 : 54



3/30



25 min



Hint

Q : Which one of the following pair is isoelectronic

A

Si, CO₂

B

H₂O, Ne

C

Na, K

D

NH₃, Mg

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24 : 52



4/30



25 min



Hint

Q : Which one has least number of isotopes

A

Tin

B

Cadmium

C

Hydrogen

D

Palladium

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24 : 50



5/30



25 min



Hint

Q : The number of moles of CH_4 which contains 3.0g of Carbon



1.0



0.75



0.5



0.25

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24 : 48



6/30



25 min



Hint

Q : A pair of compounds that has same empirical formula

A

Acetic acid and glucose

B

Acetic acid and formic acid

C

Formic acid and sucrose

D

Both a and b

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24 : 46



7/30



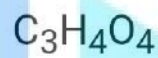
25 min



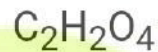
Hint

Q : An acid with molecular mass 104 contain 34.6% C, 3.85% H and rest is O the molecular formula of acid is

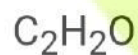
A



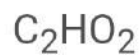
B



C



D



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24 : 42



8/30



25 min



Hint

Q : CH_2O is the empirical formula of



$\text{C}_{12}\text{H}_{22}\text{O}_{11}$



$\text{CH}_2(\text{OH})\text{CH}_2(\text{OH})$



$\text{CH}_3 \cdot \text{CH}(\text{OH}) \cdot \text{COOH}$



CH_3CHO

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9

10

11

12

24 : 40



9/30



25 min



Hint

Q : The simplest formula of a compound containing 50% of element X (At.wt = 10) and 50% of element Y (At. wt = 20) is

A

XY

B

XY₂

C

X₂Y

D

X₂Y₃

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11

12

24 : 38



10/30



25 min



Hint

Q : 3.0 g of NO gas occupies volume

A

22.424 dm³

B

2.2414 dm³

C

11.2 dm³

D

1.2 dm³

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12

24 : 36



11/30



25 min



Hint

Q : One gram molecule of different gases have all the following properties same at STP except

A

Molecules

B

Moles

C

Volume

D

Masses

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12

24 : 35



12/30



25 min



Hint

Q : 720g of water contains how many moles of water?



20



60



40



80

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6

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11

12

24 : 31



13/30



25 min



Hint

Q : An atom of carbon is twelve times heavier than ____ atom



H



Ne



He



Li

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13

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19

24 : 29



14/30



25 min



Hint

Q : What is the volume (in dm^3) of CO_2 liberated at STP, when 53gram of sodium carbonate (mol. mass = 106) is treated with excess dilute HCl in following reaction



11.2



0.448



22.414



5.51

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24 : 27



15/30



25 min



Hint

Q : 500 ml of NH_3 contains 6.00×10^{23} molecules at S.T.P. How many molecules are present in 100 ml of CO_2 at S.T.P?



6×10^{23}



1.5×10^{23}



1.2×10^{23}



3.01×10^{23}

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24 : 25



16/30



25 min



Hint

Q : Avogadro's number is the number of molecules present in

A

11.2 dm³ at STP

B

22400 cm³ at STP

C

1000 cm³ at STP

D

1 cm³ at STP

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24 : 24



17/30



25 min



Hint

Q : What volume is occupied by a mixture of 0.5g H_2 , 16g O_2 and 7.0g N_2

A

2.24 dm^3

B

22.4 dm^3

C

0.224 dm^3

D

11.2 dm^3

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24 : 21



18/30



25 min



Hint

Q : 1 g. atom of nitrogen at STP represents

A

6.02×10^{23} N₂ molecules

B

22.4 dm³ of N₂

C

11.2 dm³ of N₂

D

28 g of N₂

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24 : 19



19/30



25 min



Hint

Q :

2.8 g of an unknown gas occupies 2.24 dm³ volume at standard temperature and pressure. The gas may be

A

Carbon dioxide

B

Carbon monoxide

C

Oxygen

D

Sulphur dioxide

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24 : 11



20/30



25 min



Hint

Q : The weight of 11.2 dm^3 of CO_2 at STP would be

A

88 g

B

44 g

C

32 g

D

22 g

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24 : 09



21/30



25 min



Hint

Q : The number of moles of CO_2 which contains 16g of oxygen



0.25



1.50



0.50



1.00

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24 : 07



22/30



25 min



Hint

Q :

Efficiency of a chemical reaction is

A

Actual yield

B

Theoretical yield

C

Percentage yield

D

All of these

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24 : 05



23/30



25 min



Hint

Q :

Which of the following are limitations of chemical equations



They do not tell about the conditions of reactions



Rate of reaction



Phase change involved



All of these

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24 : 02



24/30



25 min



Hint

Q :

A chemist is more interested about _____ to express the efficiency of a chemical process

A

Theoretical yield

B

Actual Yield

C

%age yield

D

Non-limiting reactant

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24

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24 : 00



25/30



25 min



Hint

Q : Find the mass of iron which will be converted into its oxide by the action of 18g of steam as

A

42 g

B

56 g

C

4.2 g

D

28 g

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24

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23 : 58



26/30



25 min



Hint

Q : Unit of concentration for representing the trace amount of substance is

A

Molarity

B

Mole fraction

C

Molality

D

Parts per million

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23 : 55



27/30



25 min



Hint

Q : When we dissolve 15.8g of KMnO_4 in 1000cm^3 of solution then the molarity of the solution is

A

0.1M

B

0.2M

C

0.01M

D

0.02M

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30

23 : 53



28/30



25 min



Hint

Q : The sum of mole fractions of components of a solution is always equal to

A

Zero

B

One

C

Two

D

100

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24

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29

30

23 : 51



29/30



25 min



Hint

Q : The molarity of an aqueous solution of NaOH containing 8 gm in 2 litre of solution is

A

0.1 M

B

0.2 M

C

0.25 M

D

0.15 M

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29

30

23 : 48



30/30



25 min



Hint

Q : A solution of 92g of ethanol, 96g methanol and 90g water has mole fraction of ethanol equal to



0.1



0.4



0.2



0.5

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Unit-Wise - Test 1

Chemistry - Fundamental Concept

- | | | |
|------|------|------|
| 1-D | 11-D | 21-C |
| 2-D | 12-C | 22-C |
| 3-B | 13-A | 23-D |
| 4-C | 14-A | 24-C |
| 5-D | 15-C | 25-B |
| 6-A | 16-B | 26-D |
| 7-A | 17-B | 27-A |
| 8-C | 18-C | 28-B |
| 9-C | 19-B | 29-A |
| 10-B | 20-D | 30-C |

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